

Comments:

This paper deals with biodiesel from cashew nut oil. Unfortunately, there are numerous problems which make the paper not publishable.

The relevant properties of the diesel fuel used for blending are not given. What was the source of the diesel fuel? **The properties of the diesel used for blending have been included.**

Line 32. What are cardinals??? This does not make sense. **Expunged**

Section 2.3. Methoxides (alkoxides) are now more commonly for biodiesel production as they yield a higher quality product. Also, according to Table 3, the acid value of the oil was too high for direct transesterification, rather acid pretreatment would have been required. **Pretreatment was done and has now been reported.**

Section 3.0. As cashews are edible, the statement that all cashew oil could be used for biodiesel production is unrealistic. **Sometime there are surpluses or spoilt oil which can be converted to biodiesel.**

Section 3.6, line 164. The "octane" number should have been "cetane" number. Experimental cetane number determination should have been carried out. **This has been corrected.**

Section 3.11. Why should the oxidative stability of the diesel fuel be lower than that of the biodiesel. Usually one would expect the opposite. **On the contrary diesel is more stable than biodiesel.**

Section 3.12, and Table 3. Why should the acid value (value) be the same for B100 and B10 but increase for B20? This does not make sense. **This is one of the amazing results of the research. It appears this is a threshold when the acid value starts to increase with increasing blending. A good reference is Bello et al. (2011).**

Section 3.16. Determining the pH value does not make sense. **This aspect of the paper has been expunged.**

Table 1. This should not be a table. Rather give in text form in the experimental section. **This has been presented in text form.**

Table 2. There are significant differences in the amounts of some fatty acids in the profiles for the oil and B100. Why? **There are significant differences in the fatty acid profile of B100, B20 and B10 compared with CNO because transesterification chemically breaks down the molecules of the vegetable oil during which the glycerol from the triglycerides are removed and replaced with radicals from the alcohol used which changes the structure and hence the properties of transesterified oils (Bello, et al., 2011).**

Table 3.

At which temperature were the viscosity values determined? **At 40°C**

- Dynamic viscosity increases from B10 to B20 but kinematic viscosity decreases. This does not make sense. **Thanks it was a typographical error and has been corrected.**
- Biodiesel standards do not call for "calculated" cetane number. Also, the biodiesel standards ASTM D6751 prescribes the use of ASTM D 613, alternatively D6890, not D4737. **Thanks Noted.**
- The iodine value should decrease from B10 to B20, not increase as is shown here. **Thanks. It was a mix up and has been corrected**
- Giving the pH value does not make sense. **The result has been expunged**

Table 4. The column for CSO is non-sense. For example, why should CSO have such high free glycerol but low total glycerol? **The typographical error has been corrected**

The manuscript would have required some improvement of language. **Grammar has been improved on**

2nd time correction:

Section 3.12, and Table 3. Why should the acid value be the same for B100 and B10 but increase for B20? This does not make sense. **This is one of the amazing results of the research. It appears this is a threshold when the acid value starts to increase with increasing blending. A good reference is Bello et al. (2011)** **If the reviewer is not convinced this aspect of the work will be withdrawn. Rubber nut oil and egunsi melon oil follow the same trend for tests made.**

Section 3.16. Determining the pH value does not make sense. **The result has been expunged.**