



SDI Review Form 1.6

Journal Name:	Asian Journal of Chemical Sciences
Manuscript Number:	Ms_AJOCS_39216
Title of the Manuscript:	Solubility of Benzoic Acid in Six Alcohols within (288.15 to 336.15) K and Thermodynamic Properties in the Dissolution Process
Type of the Article	

General guideline for Peer Review process:

This journal's peer review policy states that **NO** manuscript should be rejected only on the basis of '**lack of Novelty**', provided the manuscript is scientifically robust and technically sound. To know the complete guideline for Peer Review process, reviewers are requested to visit this link:

(<http://www.sciencedomain.org/page.php?id=sdi-general-editorial-policy#Peer-Review-Guideline>)



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PART 1: Review Comments

	Reviewer's comment	Author's comment (if agreed with reviewer, correct the manuscript and highlight that part in the manuscript. It is mandatory that authors should write his/her feedback here)
Compulsory REVISION comments	(1) The valid number should be consistent for one variable. Please check table 3.	The valid number in Table 3 has been revised in order to be consistent for one variable.
Minor REVISION comments	(1) page 3: It is not clear who did the GC and HPLC analysis in the Table 1 (supplier or authors). If the supplier, it should be clearly indicated in Table 1. If the authors, the corresponding methods should be briefly described in the manuscript. (2) page 6: The correlation models used are well known. Please make the model introduction concise and refer to literatures as much as possible. (3) Really it will be interesting to see activity coefficients calculated by activity coefficient models. I suppose the Table with this information will be very useful.	(1) The GC and HPLC analysis in the Table are done by supplier and it has been added clearly in Table 1. (2) The correlation models in sections 3.1.1, 3.1.2, 3.1.3 has been minimized (3) The regression parameters of the activity coefficient models were listed in Table 3. And according to these parameters in equation 4 and equation 5, activity coefficients can be calculated easily. But considering the limited space, the data were not listed in Table.
Optional/General comments	(1) Page 4: Why did the author use gravimetric method in this study? Any advantage or consideration? (2) Page 4: The authors must clarify as how they have minimized the error accompanied with the vaporization of solvent during the solute was added and the equilibration period.	(1) The advantage of the gravimetric method is that it is easy to operate. Therefore, it reduces the source of the experimental error. By changing the system temperature, it can continuously measure the solubility data at different temperatures with less time and high efficiency. (2) A condenser was equipped and all ports tightly sealed to avoid the vaporization of the solvent..