

A Study of Potential Corrosivity of Borehole Waters in Eket, Nigeria by Saturation Index Measurement

Keywords : Calcium Carbonate saturation index, saturation p H, corrosion, Borehole waters

Abstract

A study of water chemistry in some boreholes in Eket, Nigeria, was conducted to determine the calcium carbonate saturation index, in order to establish the corrosion tendency of the waters towards metal substrates in piping systems, boilers, heat exchangers and submersible pumps. Calcium carbonate saturation index using Langelier system was determined to be more negative (- 4.16 to - 5.2) below the optimum range of - 0.5 to + 0.5 for water that is in equilibrium with calcium carbonate. The water saturation p H (pHs) determined for the waters sampled in the area had values that ranged from 8.82 to 9.55, with mean value of 9.09. The values were far above the pH values of the waters from all the locations sampled, which ranged from 4.12 to 5.34, with a mean pH of 4.48 . The water chemistry was of same characteristic for all the studied locations, indicating that the waters were drawn from same regional aquifer, have same corrosive tendency and the possibility of metal substrates to be attacked by water used for both industrial and domestic purposes in the study area. The more negative the calcium carbonate saturation index, the greater the possibility of the water in its corrosive action especially in oxygenated waters. Therefore, the waters in Eket are likely to exhibit remarkable corrosive tendency against metal substrates. The implication of this finding is that in the study area, there has to be a strategic corrosion control program in piping systems and industrial equipment such as heat exchangers, boilers, domestic water heaters and submersible pumps in tackling corrosion challenges.

1.0 Introduction

One of the most common problems affecting domestic water supplies is corrosion, a chemical process that slowly dissolves metal, resulting in deterioration and failure of plumbing pipes, fixtures and water-using equipment (Garone et al, 2012). Corrosion attacks the metal substrate and gradually thins the entire metal surface, often causing red-colored stains in iron or steel plumbing systems or blue-green stains in copper and brass plumbing systems (Etesin

et al, 2012 ; Oram, 2001). Another type of corrosion attacks are small areas where deep pits can develop and penetrate pipe or tank walls. This type of corrosion may not add significant amounts of iron or copper to the water, but can create small holes in a pipe or tank that destroy its usefulness, cause water leaks, and result in major water damage to a home or business (APHA, AWCF, APWA, 1995). In another case, corrosion may cause oxidation of metals, which involves conversion of copper or other base metal to an oxidized form in a process similar to the rusting of steel (Hermanson, 1991). It often results in reduced water flow through supply lines and destruction of water valves and other machined water flow control surfaces, thus resulting in internal and external leaks at valves and faucets.

Corrosion is a complex series of reactions between the water metal surfaces and minerals in which the water is stored or transported. With respect to the corrosion potential of drinking and process water, the primary concerns include potential presence of toxic metals, such as lead and copper, deterioration of and damage to the household plumbing and aesthetic problems such as stained laundry, bitter taste and greenish-blue stains around basins and drains (Oram, 2001).

Nearly all metals will corrode to some degree. The rate and extend of the corrosion depends on the degree of physical and chemical characteristics of the media, metal and environment. In soft water, corrosion occurs because of lack of dissolved cations, such as calcium and magnesium in the water. In water that forms scales, a precipitate of calcium or magnesium carbonate forms on the inside of the piping system (Etesin et al, 2012). This scale can inhibit the corrosion of the metal substrate, because it acts as a barrier, but can also clog piping system. However, high concentrations of sodium, chloride, or other ions will increase the conductivity of the water and promoting corrosion (Oram, 2001).

Water quality is very critical in influencing the operating efficiency and life of equipment, including pumps, piping systems and wells. Therefore, acidic water with a p H value of less than 7 is usually corrosive while, alkaline water with pH less than 7 is less corrosive (GWM, USA, 1985).

Corrosive water can be defined as a condition of water quality which will dissolve metals from metallic plumbing at an excessive rate. Some of the factors that makes water corrosive include, low pH (below 7), a lower alkalinity value (below 35 mg/l CaCO_3), a high specific conductivity as a measure of total dissolved solid (TDS) and high temperature (Environmental Facts Sheet, 2009).

63 An indicator of water corrosiveness that is widely in use is calcium carbonate (CaCO_3)
 64 saturation index (SI), also referred to as calcium carbonate solubility equilibrium, and is
 65 highly important in many phases of water chemistry (Bryan, 2009). It is used to evaluate the
 66 scale-forming and scale-dissolving tendencies of water.

67 Several indices are used to describe the behavior of calcium carbonate in water, oil, or gas
 68 mixtures. In 1936 Wilfred Langelier developed a method for predicting pH at which water is
 69 saturated in calcium carbonate, which is called saturation pH (pHs), (Egereonu *et al.*, 2012).

70 Langelier saturation index (LSI) is the difference between the actual system pH and the
 71 saturation pH (pHs). pH is a common measure of water quality that is often related to
 72 corrosion of plumbing system components like lead and copper (APHA, AWWA, APWA, 1995).
 73 Lead is a dangerous metal with many health effects, especially in children.
 74 Lead usually occurs in drinking water supplies from corrosion of lead solder or lead
 75 impurities in plumbing fixtures (Bryan, 2009).

76 The assessment of scale-forming and scale-dissolving tendencies of water is useful in
 77 corrosion control program and also in preventing calcium carbonate scaling in piping, and
 78 equipment such as industrial heat exchangers, boiler feed water heaters and domestic water
 79 heaters, including tea kettle (Etesin *et al.*, 2012). In this consideration, there are three classes
 80 of water that can be defined by calcium saturation index namely, waters oversaturated with
 81 CaCO_3 , which tend to precipitate CaCO_3 scale; waters under saturated with CaCO_3 , which
 82 tend to dissolve calcium carbonate and waters saturated with CaCO_3 , which are in
 83 equilibrium with calcium carbonate and have neither CaCO_3 precipitate nor dissolving
 84 tendencies (APHA, AWWA, APWA, 1995).

85 The aim of this study is to estimate the calcium carbonate saturation index of some boreholes
 86 in Eket and also evaluate the influence of water quality parameters such as calcium, total
 87 dissolved solids, alkalinity, pH, temperature and conductivity on the calcium carbonate
 88 solubility equilibrium and to establish the corrosion tendency of water in the study area.
 89 Companies in the area use water from this source as process and portable water, the local
 90 communities also use water from this source for both portable and domestic uses. The huge
 91 costs of corrosion control industrial and domestic equipment in the area lend some impetus to
 92 this study.

93 The study will provide adequate data on potable water supply in the study area, whereby such
 94 data will be useful in evaluation of the extent of potential corrosivity of water. The data to be

95 generated from the study will help in the future water resources planning development in the
96 study area.

97 The Langelier saturation index (LSI) is a formula developed from studies conducted by Dr.
98 Wilfred Langelier in the early 20th century. The LSI is the basis for water balance and
99 saturation (APHA. AWCF, APWA, 2005)

100 The Langelier saturation index (LSI) is an equilibrium model derived from the theoretical
101 concept of saturation and provides an indication of the degree of saturation of water with
102 respect to calcium carbonate. It can be shown that the Langelier saturation index (LSI)
103 approximates the base 10 logarithm of the calcite saturation level (APHA. AWCF, APWA,
104 2005).

105 The Langelier saturation level approaches the concept of saturation using pH as a main
106 variable. The LSI can be interpreted as the pH change required ,bringing water to equilibrium
107 (Kriesi, 2017).

- 108 • If LSI is negative: no potential to scale, the water will dissolve CaCO_3
- 109 • If LSI is positive: scale can be formed and CaCO_3 precipitation will occur
- 110 • If LSI is close to zero: borderline scale potential.

111 The LSI is probably the most widely used indicator of cooling water scale potential. It is
112 purely an equilibrium index and deals only with the thermodynamic driving force for
113 calcium carbonate scale formation and growth.

114 It provides no indication on how much scale or calcium carbonate will actually precipitate to
115 bring water to equilibrium. It simply indicates the driving force for scale formation and
116 growth in terms of pH as a master variable. In order to calculate the LSI, it is necessary to
117 know the alkalinity (mg/l Ca^{2+} as CaCO_3), total dissolved solids (mg/l TDS), the actual pH,
118 and the temperature of the water ($^{\circ}\text{C}$).

119 LSI can be defined as (APHA, AWCF,APWA 2005) :

$$120 \text{ LSI} = \text{pH} - \text{pHs}$$

121 Where:

- 122 • pH is the measured water pH
- 123 • pHs is the pH at saturation in calcite or calcium carbonate and is expressed as :

$$124 \text{ pHs} = (9.3 + A + B) - (C + D)$$

125 Where:

- 126 • $A = (\log_{10} [\text{TDS}] - 1) / 10$
- 127 • $B = - 13.12 \log_{10} (^{\circ} \text{C} + 273) + 34.55$
- 128 • $C = \log^{10} [\text{Ca}^{2+} \text{ as CaCO}_3] - 0.4$
- 129 • $D = \log_{10} [\text{alkalinity as CaCO}_3]$

130 The consequences of corrosion are many and varied and the effects of these on the safe,
 131 reliable and efficient operation of equipment or structures are often more serious than the
 132 simple loss of a mass of metal (Corrosion Report, 2001) . Failures of various kinds and the
 133 need for expensive replacements may occur even though the amount of metal destroyed is
 134 quite small.

135 Corrosion is a natural process that occurs when metals are in contact with oxygen and react to
 136 form metal oxides. All water is corrosive to some degree as it contains some amount of
 137 dissolved oxygen. The rate of corrosion depends on a number of factors including acidity or
 138 low pH, electrical conductivity, oxygen concentration and water temperature (Oram, 2001).

139 In addition to corrosion, dissolution of metals occurs when the water is extremely low in
 140 dissolved salts or in the presence of certain water-borne ions. All materials have a particular
 141 level of solubility and in the case of corroded plumbing, the concentration of the copper or
 142 other plumbing material metal is lower in the water than that material's solubility. As a
 143 result, the plumbing material is gradually dissolved. While this process is usually very slow,
 144 certain water-borne ions can react with and bind the recently dissolved metal allowing more
 145 rapid loss. While corrosion and dissolution are fundamentally different, the end result is
 146 similar and so both are often discussed together under the general term corrosion (Subyami,
 147 2004)

148 Ground water can be acidic or alkaline in pH depending on several factors. Rainfall is
 149 typically acidic because it picks up carbon dioxide as it falls to the earth, forming carbonic
 150 acid. As water percolates through the soil, it also can come in contact with other acid-
 151 forming materials such as decaying organic matter (Obiefuna and Orazulike, 2011). In areas
 152 where underground strata contain limestone or dolomite, the acid is neutralized and the water
 153 is usually alkaline and hard with pH values between 7 and 8. Where limestone or dolomite is
 154 not present underground, the percolating water will retain its acidity and groundwater will
 155 typically have pH values between 6 and 7 (Edet and Ekpo, 2008).

Corrosion is more likely and more rapid at higher water temperatures. The rate of corrosion increases by a factor of three to four as water temperature rises from 60°F to 140°F. Above 140°F, the rate of corrosion doubles for every 20°F increase in water temperature (McFarland, *et.al.*, 2012).

2.0 The Study Area

Eket Local Government Area is one of the local government areas in Akwa Ibom State , Nigeria, playing host to a major multinational oil company, Exxon Mobil. Eket is located between latitude 4°31¹ and 4°45¹ North and between longitude 7°52¹ and 8°02¹ East (Figure1). The study area is underlain by sedimentary formation of Late Tertiary and Holocene ages (Magnus et al, 2012) . Deposits of recent alluvium and beach ridge sands occur along the coast and the estuaries of the Imo and Qua Iboe Rivers and also along flood plains of creeks. The study area consists of mainly of coastal plain sands. The sands are mature, coarse and moderately sorted. The coastal plain sands, otherwise known as the Benin formation overlies the Bende-Ameki formation and dips south westward (.Akankpo and Igbokwe, 2011 ; Mbonu and Ebeniro, 1991).. The landscape of Eket comprises of a generally low-lying plain and riverine areas with no portion exceeding 175 metres above mean sea level. The physiography of Eket is that of a beach ridge complex characterized by a succession of sub parallel sand ridges (Figure 1).

The physical relief of Eket Local Government Area is basically flat, though with some marshy river-washed soils around the banks of Qua Iboe River (ALSCON, 1997) . Eket Local Government falls within the tropical zone wherein its dominant vegetation is the green foliage of trees/shrubs and the oil palm tree belt. The Local Government has two seasons: the wet season and the dry season. Eket has abundant deposit of crude oil and clay. Forest resources include timber, palm produce while the area is also noted for seafood production. Farm crops range from yam, cassava, cocoyam, plantain to maize and vegetables (Akankpo and Igbokwe , 2011).

3.0 Materials and Methods

Water samples from boreholes were collected from selected villages within Eket L.G.A : (a) No.12 Ekpeyu street (S1); (b) Ofonmbuk street(S2) ; (c) Hospital road(S3); (d) Ikot Ibiok(S4); (e) Iko Ewak(S5); (f) Liverpool(S6); (g) Iquita street(S7), (h) Etebi street(S8), (i) Nung Uso street(S9) ; (j) Ikot Ebok(S10) . Prior to the collection of water samples, the nozzles of the tap were flamed and sterilized by cleaning with cotton wool soaked in a

188 methylated spirit to avoid contamination. The tap were allowed to run for 5-10 minutes to
189 bring forth the underground water and not the one in the distributing pipeline.

190 Water samples were collected in a 500ml polyethylene bottles and stored under ice in coolers
191 before being transported to the laboratory for analysis within two hours. Such physico-
192 chemical parameters like pH, conductivity, total dissolved solids (TDS) and temperature were
193 determined at the field using a multimeter (DDJ-303A). Calcium content of the water was
194 determined by Flame Atomic Absorption Spectrometer (Unicam 939/959) using nitrous oxide
195 and acetylene flame, at a wavelength of 422.7nm (Welz, 1985).

196 The alkalinity of water was determined by titrimetric method, in accordance with the method
197 of APHA. AWCF, APWA (2005). In this study, since all the water samples analyzed had pH
198 below 8.3, only total alkalinity was determined by titration using dilute tetraoxosulphate (VI)
199 acid (APHA. AWCF, APWA ,2005).

200 The saturation index is a calculated number used to predict the calcium carbonate stability of
201 water from surface or underground bodies like boreholes and wells. In order to estimate the
202 calcium carbonate saturation index, several water quality characteristics must be measured
203 such as alkalinity, total calcium, pH, temperature, conductivity, total dissolved solid, ionic
204 strength etc (APHA, AWCF, APWA, 1995).

205 Various equations have been proposed for calculating the calcium carbonate saturation index,
206 among which are, Tillmans and Heublein (1912); Greenfield and Baker (1920); Johnson and
207 Fear (1929); Langelier and others (1936). Langelier developed a method for predicting the
208 pH at which water is saturated in calcium carbonate (pH_s), whereby saturation index is
209 expressed as the difference between the actual pH and the saturated pH (Demartini, 1938). If
210 the actual pH of water is below the calculated saturated pH, the saturation index is negative
211 and the water a very limited scaling potential. If however, the actual pH exceed pH_s the
212 saturation index is positive and is being super saturated with calcium carbonate, the water has
213 the tendency to form scale. At increasing positive index values, the scaling potential
214 increases.

215 The major task in saturation index estimation is the determination of the saturation pH (pH_s),
216 which is expressed as :

$$217 \quad pH_s = PK_2 - PK_s + p[Ca^{2+}] + p[HCO_3^-] + 5p_{fm}$$

218 Where:

219 K_2 is second dissociation constant for carbonic acid at the water temperature

220 K_s is the solubility product constant for CaCO_3 at the water temperature

221 $[\text{Ca}^{2+}]$ is the concentration of calcium ion in g-mole per litre

222 $[\text{HCO}_3^-]$ is the concentration of bicarbonate ion in g-mole per litre and

223 f_m is the activity coefficient for monovalent ionic species.

224 In aqueous systems, different isomorphs of calcium carbonate can form like calcite, aragonite
225 and valerite, with different solubility properties. Depending on the isomorph of calcium
226 carbonate in the aqueous system, values of PK_2 , PK_s and $\text{p}f_m$ required to solve equation are
227 available in literature (APHA, AWCf, APWA, 1995). Alkalinity in water or waste water is as
228 the result of the presence of hydroxide, carbonate and bicarbonate ions. The alkalinity of
229 water is a measure of its capacity to neutralize acids. However, bicarbonate is the major
230 contributor to alkalinity and it arises from the action of carbon dioxide in percolating water
231 on basic minerals in soil and rocks (APHA, AWCf, APWA, 2005; Etesin et al, 2012).

232 **4.0 Results and Discussion**

233 The results of the analyses of the borehole waters from the study area and calculations are
234 presented in Tables 1, 2, and 3, while Figures 2, 3, 3, 4 and 5 are plots of the variations in the
235 physicochemical properties of the waters and the calcium carbonate saturation index of the
236 waters.

237 The temperature of the waters ranged from 21.5°C to 27.2°C with a mean of $25.05^\circ\text{C} \pm$
238 1.17°C . The pH of waters had a range of 4.12 to 4.72, with a mean of 4.48 ± 0.37 . For all the
239 waters the pH indicated acidic range (Table 1). The pH values were outside the limits of the
240 World Health Organization (WHO, 2016) , of 6.5 to 8.5 for potable water.

241 The Calcium concentration of the waters ranged from 3.61 mg/litre to 20.02 mg/litre, with a
242 mean of $11.55 \text{ mg/litre} \pm 5.55 \text{ mg/litre}$ (Table 1). The conductivity of the waters had a range
243 of $51.43 \mu\text{S/cm}$ to $170.8 \mu\text{S/cm}$, with a mean of $117.58 \mu\text{S/cm}$. The high conductivity recorded
244 at some stations could be attributed to salt water intrusion as the stations are close to the
245 coastal line, whereby water with high levels of sodium, chloride or other ions will increase
246 the conductivity of the water and promoting corrosion (Oram, 2001).

247 The total dissolved solid (TDS) of the waters had a range of 26.1 mg/litre to 85.2 mg/litre,
248 with a mean of $59.09 \pm 4.09 \text{ mg/litre}$. The range of alkalinity recorded for the waters was
249 49.98 mg/litre to 75.1 mg/litre, with a mean alkalinity of $58.72 \text{ mg/litre} \pm 10.57 \text{ mg/litre}$

(Table 1). The range of the total dissolved solids (TDS) obtained for waters in the study area were far below the maximum 1000 mg/l TDS limit of the World Health Organisation (WHO, 2006) , indicating that the waters are fresh (Obiefuna and Orazulike, 2011).

The range of saturation p H (pHs), at which water is saturated with calcium carbonate determined for the waters was 8.82 to 9.55, with a mean of 9.09 ± 0.23 (Table 2). The saturation index calculated for all the waters studied from the area had values that ranged from -5.2 to -4.16 , with mean value of -4.62 ± 0.37 (Table 3), compared to the optimum Langelier saturation index of -0.5 to $+0.5$ (APHA, AWCF, APHA, 2005) . The mean saturation index for the waters was out of the optimum range of the Langelier saturation index, indicating that all the waters studied in the study area are calcium carbonate dissolving and under saturated with calcium carbonate (Pisigan,1984; Etesin et al, 2012).

The negative values of calcium carbonate saturation index determined for the waters in this study, indicated that calcium carbonate will not precipitate in the waters from the study area , hence there may be a possibility of corrosion of metal substrates in contact with the waters, especially if dissolved oxygen is high (MTI, 1995). The relationship between the saturation p H (pHs) and Langelier saturation index is illustrated in Fig. 4.

The Langelier saturation index mean value obtained in this study is -4.62 ± 0.37 , which is lower than the value of -2.89 reported for waters in Ikot Abasi and the mean saturation p H (pHs) obtained in this study is 9.09 ± 0.23 , which is higher than the value of 8.93 reported for waters from Ikot Abasi (Etesin et al, 2012).

In a similar study reported on the geochemical studies of groundwater systems of semiarid Yola area, Northeast, Nigeria (Obiefuna and Orazulike, 2011), the saturation index of the waters were less than unity , as also obtained in this study, which indicated waters undersaturated with respect to calcite , dolomite and gypsum (Etesin et al, 2012; APHA, AWCF, APWA, 2005) . The unsaturation of calcite, dolomite and gypsum in the waters suggests that their soluble ions Calcium, magnesium ,sulphate and carbonate components are limited by mineral equilibrium, and the water has the potential to be corrosive , mostly to metal substrates (Environmental Fact Sheet, 2009). Mineral dissolution in groundwater is affected by the presence of carbon dioxide and temperature of the underground water (Freeze and Cherry, 1979). In the study area, the temperature of the waters ranged from 21.5°C to 27.2°C with a mean of $25.05^{\circ}\text{C} \pm 1.17^{\circ}\text{C}$, while alkalinity of the waters ranged from 49.98 mg/l to 75.10 mg/l , with the mean of 58.72 mg/l , which is less than the limits set by the

WHO (2006). The low alkalinity in the water of the study area , may likely impact on the quality of the water in terms of being corrosive.

The correlation between calcium carbonate saturation index and alkalinity of the waters from the study area was not significant using Pearson's correlation coefficient (R). Equilibrium calculations help to determine the saturation state of groundwater with respect to various minerals and gases of interest, and more so, are useful in predicting the presence of reactive minerals in the groundwater system and estimating mineral reactivity (Deutsch, 1997 ; Obiefuna and Orazulike, 2011).

5.0 Conclusion and Recommendation

5.1 Conclusion

Calcium carbonate saturation index had been determined for waters from boreholes in Eket , as calculated by Langelier saturation index. The results indicated more negative values below the optimum range of water that is in equilibrium with calcium carbonate, which is calcium carbonate dissolving without scale formation.

The water saturation pH (pHs) were far above the pH values of the waters from all the locations sampled. The water chemistry were of same characteristics indicating that the waters were drawn from similar aquifer, have same corrosive tendency and the possibility of metal substrate to be corroded by water used for both industrial and domestic purposes in the study area. The more negative the calcium carbonate saturation index, the more notorious the water is in its corrosive action (Larson and Budswell, 1943; Larson and Skold, 1958).

Therefore, the waters in Eket are likely to exhibit some corrosive tendencies against metal substrates. The implication of this finding is that there has to be a strategic corrosion control program.

As indicated, corrosion is a major problem that impacts every industry (Radia, 2004). In the study area, just as in other environments, attention to corrosion in the materials selection, construction, operation and maintenance of facilities employing corrosion - able metals will save billions of dollars in repair/maintenance and replacement costs. To save the huge cost of corrosion, every attempt must ensure the breaking of the linkage in the triangles of corrosion namely anode, cathode and electrolyte. And one of such preventive and controlling measures will be coating the surface of the metal with nonconductive organic coatings like paint and primers.

5.2 Recommendation

For the boreholes in the study area, as low cost is a significant part of the decision process, one major consideration will be the addition of a layer of calcite (sanitary calcium carbonate) to the bottom of the well, that will dissolve slowly neutralizing the water's corrosiveness. On the other hand for high technology approach, it is advisable to add a dilute solution of sodium carbonate or bicarbonate to the water treatment system as to raise the pH and alkalinity in order to combat corrosiveness of the water.

For further study on the water quality of the area, it is suggested that heavy metal contents of water and the ground water dynamic using geochemical modeling be undertaken.

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 414 Verlagsgesellschaft.mbH(D6941). Weinheim,. Germany. 277-278
- 415 **Table1: Results of physicochemical analyses of borehole water sample from Eket,**
 416 **Nigeria**

Sampling borehole	DO (mg/l)	pH	Redox potential (mv)	Ca (mg/l)	Conductivity (µs/cm)	TDS (mg/l)	Alkalinity (mg/l)	Temperature (° C)
12 Ekpeyu street (S1)	4.11	4.69	160	16.8	154	80.3	49.98	21.5
Ofonmbuk street (S2)	3.32	4.42	196	15.4	167	83.2	55.1	26
Hospital Road (S3)	4.23	4.58	153	15.4	130	65.1	75.1	24.4
Ikot Ibiok (S4)	4.42	4.23	113	8.1	51.43	26.1	50.01	27.2
Iko Ewak(S5)	3.06	5.34	40	3.61	62.7	30.1	49.98	26.5
Liverpool(S6)	3.60	4.12	113	8.06	102.7	52.0	70.01	26.3
Iquita street(S7)	3.45	4.14	110	4.12	63.1	31.6	72.01	26.3
Etebi street(S8)	5.43	4.32	165	10.4	170.8	85.2	50.01	24.0
Nung Uso street(S9)	4.58	4.72	75.2	20.02	150.6	75.3	50.03	24.3
Ikot Ebok(S10)	4.80	4.25	162	13.6	123.5	62	64.98	24
Mean	4.1	4.48	128.72	11.55	117.58	59.09	58.72	25.05
Standard	1.06	0.37	47.26	5.55	45.29	22.99	10.57	1.72

deviation								
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418 **Table2: Calculated pHs values of the sampled waters**

Sampling borehole	A	B	C	D	pHs
12 Ekpeyu street (S1)	0.090	2.156	0.825	1.699	9.02
Ofonmbuk street (S2)	0.092	2.069	0.788	1.741	8.93
Hospital Road(S3)	0.081	2.099	0.788	1.876	8.82
Ikot Ibiok(S4)	0.042	2.046	0.508	1.699	9.18
Iko Ewak(S5)	0.048	2.059	0.158	1.699	9.55
Liverpool(S6)	0.072	2.063	0.506	1.845	9.08
Iquita street(S7)	0.049	2.063	0.215	1.858	9.34
Etebi street(S8)	0.093	2.107	0.617	1.699	9.18
Nung Uso street(S9)	0.088	2.102	0.901	1.699	8.88
Ikot Ebok(S10)	0.079	2.107	0.734	1.813	8.94
Mean	0.073	2.087	0.604	1.763	9.09
Standard deviation	0.02	0.03	0.26	0.08	0.23

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420 $\text{PHs} = (9.3 + A + B) - (C + D)$, $A = (\log_{10} [\text{TDS}] - 1)/10$, $B = -13.12 \log_{10} (T^0C + 273) + 34.55$, $C = \log_{10}$

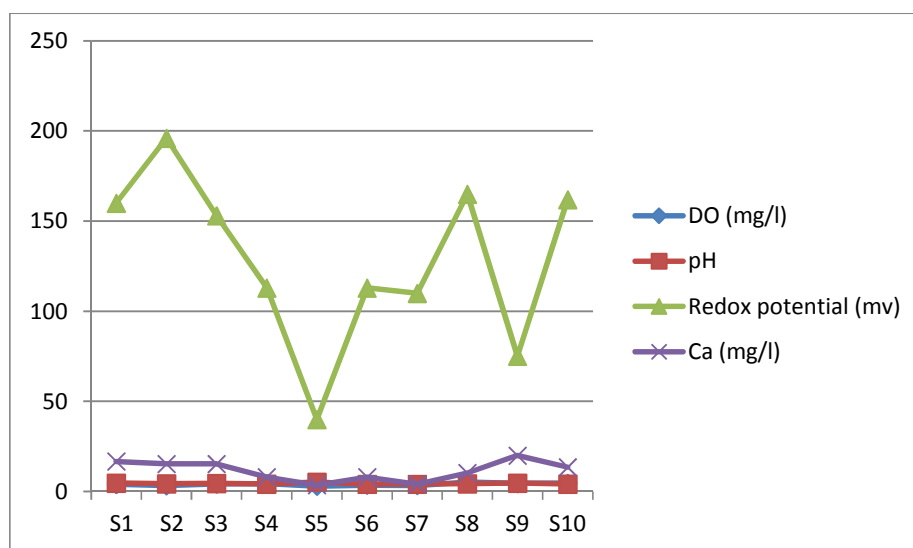
421 $[\text{Ca}] - 0.40$, $D = \log_{10} [\text{Alkalinity}]$, All the concentration are in milligram per litre.

422 **Table3: Calculated LSI values of Eket underground waters**

Sampling borehole	p H	PHs	LSI (saturation index)
12 Ekpeyu street(S1)	4.69	9.02	- 4.33
Ofonmbuk street (S2)	4.42	8.93	- 4.51
Hospital Road(S3)	4.58	8.82	- 4.28
Ikot Ibiok(S4)	4.23	9.18	- 4.95
Iko Ewak(S5)	5.34	9.55	- 4.21

Liverpool(S6)	4.12	9.08	- 4.96
Iquita street(S7)	4.14	9.34	- 5.2
Etebi street(S8)	4.32	9.18	- 4.86
Nung Uso street(S9)	4.72	8.88	- 4.16
Ikot Ebok(S10)	4.25	8.94	- 4.69
Mean	4.48	9.09	- 4.62
Standard deviation	0.37	0.23	0.37

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Fig. 2 : DO, p H , Redox Potential and Calcium of Waters from Study Area

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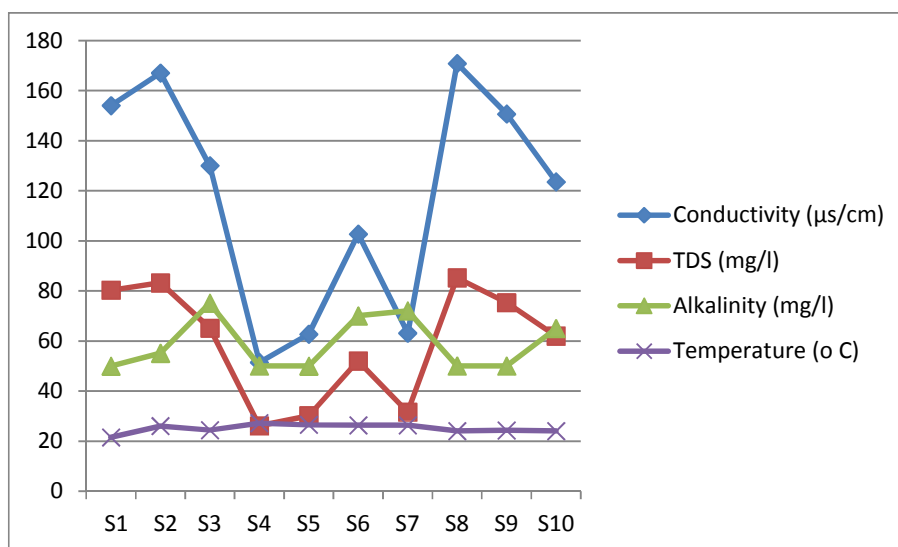


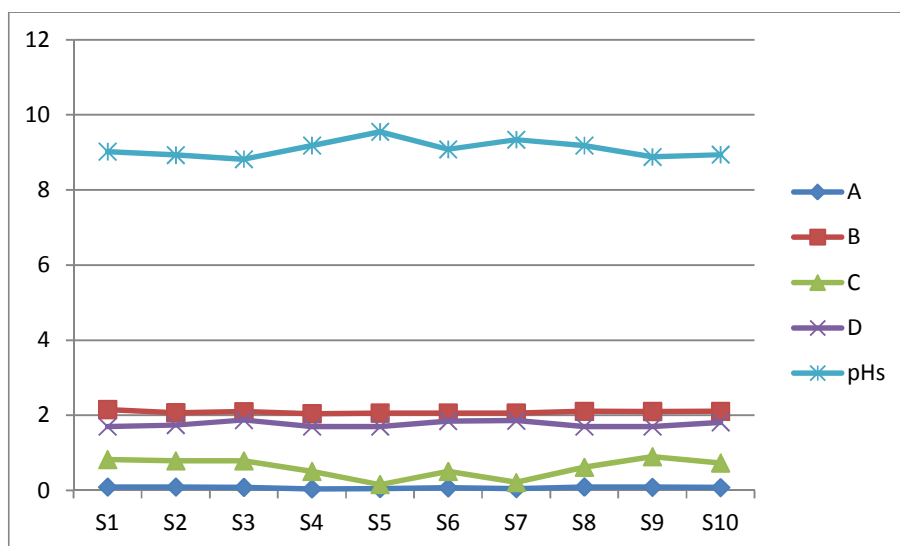
Fig. 3 : Cond, TDS , Alkalinity and Temp. of Waters from Study Area

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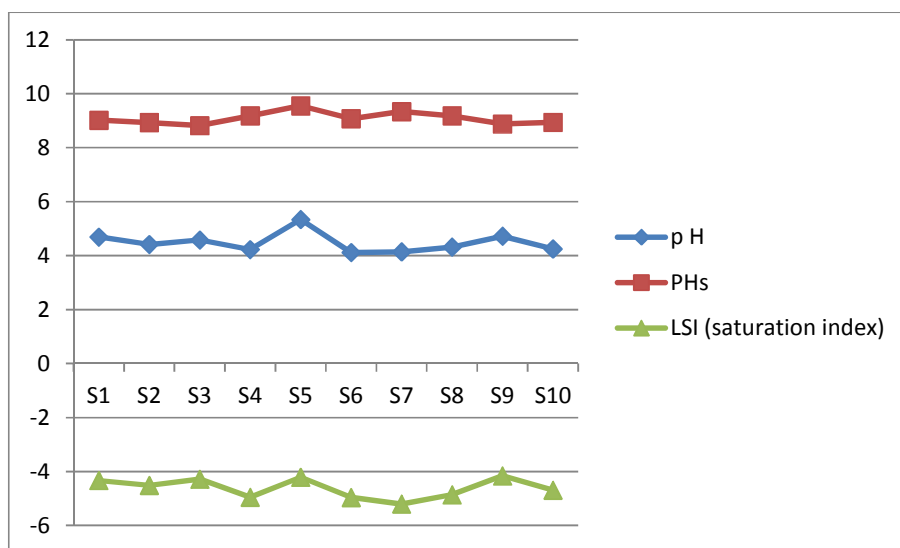
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Fig.4 : Calculated A, B, C, D and p Hs values



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Fig. 5 : p H, p Hs, and Calculated LSI of Waters from the Study Area.