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SDI FINAL EVALUATION FORM 1.1

PART 1:

Journal Name:	**Advances in Research**
Manuscript Number:	**Ms_AIR_23224**
Title of the Manuscript:	**Equilibrium Isotherm Study for Removal of Mn (II) from Aqueous Solutions by Using Novel Bioadsorbent Tinospora cordifolia**
Type of the Article	**Original Research Articles**

PART 2:

FINAL EVALUATOR'S comments on revised paper (if any)	**Authors' response to final evaluator's comments**
Line 77: After that, the biomass was stay for 30 minutes with HCl (0.1M) solution. After that, the biomass stayed inside HCL(0.1M) solution for 30 minutes. **Let the author remove was** Line 125, 134, 141, 143 and 144: You repeatedly represented the unit of mass of biosorbents as mg/l. Apply this on the formula for obtaining adsorbent capacity and use dimensional analysis to crosscheck if u will get mg/g as the unit. My analysis shows that you will get L as unit of uptake capacity instead of mg/g or g/g as the case may be as reported previously. I am not comfortable with unit of uptake capacity or adsorption capacity as adopted by the author going by his unit of mass of biosorbent which he measured in mg/l. The adsorption capacity or metal uptake rate is calculated using the following mass balance equation: $q_t = (C_0 - C_t)\frac{V}{M}$ (1) where q_t (mg/g or g/g or mg/mg); is the adsorption capacity or metal uptake rate or uptake capacity at time t C_o (mg/l), and C_t (mg/l), are the initial metal ion concentration and metal ion concentrations at time t in the solution respectively; *V*(l) is the solution volume and M (g) is the mass of biosorbent So, if unit of mass as he said is mg/l, unit of concentration becomes mg/l and the unit of volume becomes l then from equation 1 above, the unit of uptake capacity = $q_t = \left(\frac{mg}{l}\right) x \frac{l}{\frac{mg}{l}} = l$ Let the author check the dimensional analysis as shown above come up with something more convincing. Outside these I think he has done a good work.	

Reviewer Details:

Name:	**Nwidi Ignatius Chineke**
Department, University & Country	**University of Nigeria, Nsukka, Nigeria**