

Modeling of advanced photo oxidation of alizarin red s dye using TiO₂ as sensitizer

Abstract: The objective of this study was to investigate the experimental study and theoretical study of advanced photo oxidation using TiO₂ as sensitizer and to compare between these two studies. This work contains two parts. The first part deals with removing efficiency of an industrially important dye Alizarin Red S (ARS) from aqueous media using TiO₂ anatase as adsorbent. Various parameters, which can influence the degradation rate like, initial concentration, temperature, pH, adsorbent dose were investigated. The second part deals with the theoretical study of the TiO₂ surface and the pollutant substance (ARS) using the program (Spartan '08 V1.2.0). we used two kinds of calculations, the first was the Equilibrium Geometry optimization at ground state by Density Functional Theory/B3LYP method and Semi-empirical/PM3 method to predict the most stable geometrical formula for TiO₂ anatase (001),(101) and (ARS) pollutant in different medium (neutral, acidic, basic). The program showed that the method DFT/B3LYP give the best results compared with DFT/PM3 and that the basic medium was better than acidic and natural medium which means that the highest rate of (ARS) degradation was in basic medium. The second calculation was the energy geometry at ground state by density functional theory/EDF2 method to calculate standard thermodynamic quantities for (ARS) pollutant and TiO₂ anatase (001),(101) surface that showed the highest stability for them in the basic medium.

The results from theoretical study agreed with the results obtained in the experimental study. Results from the experimental study revealed that the photodegradation reaction from Alizarin Red S followed first order and exothermic reaction and optimum parameters for removal of this dye from aqueous media were (0.16 kg/l, 20ppm, temp.30°C, pH9). Theoretical study DFT/B3LYP method showed better results than DFT/PM3 to predict the most stable geometrical formula, and TiO₂ anatase (001) surface was more active than TiO₂ anatase (101) surface.

Keywords: [Molecular modeling, Alizarin Red S Degradation, TiO₂ anatase (001)(101)surface, DFT Methods, Equilibrium Geometry optimization, energy geometry]

1. INTRODUCTION

Textile dyeing process is a significant source of contamination responsible for the environmental pollution. To meet the demand of textile industry, a large variety of dyes are annually produced in tremendous quantities (7×10^5 Tones) in the world. Thus the volume of wastewater containing processed textile dyes is on steady increase⁽¹⁾. Most of the dyes used in the textiles industries are stable to light and non biodegradable⁽²⁾. In order to reduce the risk of environmental pollution from such wastewater, it is necessary to treat them before discharging into the environment⁽³⁾. Today more than 10,000 dyes have been incorporated in colour index⁽⁴⁾. Many physical, chemical, and biological methods are employed to eliminate dyes from the wastewater. Physical methods are preferred over chemical and biological methods, because firstly they are simple and inexpensive and secondly do not produce byproducts in water bodies^(5,6). Alizarin red-S is an anionic dye which are widely used in woven fabrics, wool, cotton textiles^(7,8).

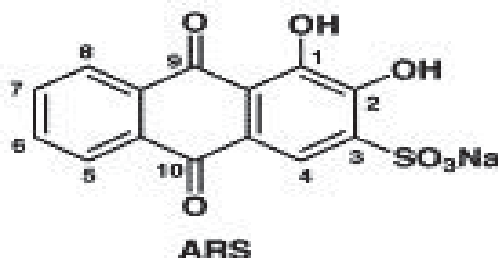


Figure .1 : Alizarin Red S

Alizarin Red S (1,2-dihydroxy-9,10-anthra-quinone sulfonic acid sodium salt, Fig. 1) is a water- soluble, widely used anthraquinone dye. It is synthesized by sulfonation of alizarin which is a natural dye obtained from madder (*Rubiatinctorum* L. Rubiaceae). Anthraquinone dyes like ARS are durable pollutants, released especially by textile industries in the aquatic ecosystems.

Removal of these dyes from industrial wastewaters is a crucial process, from both economical and environmental points of view⁽⁹⁾.

In this study TiO₂ is used as an adsorbent. K.M.Joshi et al⁽¹⁰⁾ indicated that TiO₂ is a good adsorbent for the removal of Alizarin Red S from aqueous media at pH 8. The aqueous phase photochemical degradation of dyes like Alizarin red-S has been studied in the presence of semiconducting oxide like TiO₂ and ZnO⁽¹¹⁾. In recent years, photocatalytic oxidation technology has been found capable of effectively removing pollutants⁽¹²⁾, Anatase TiO₂ is in the focus of photocatalytic research because of its strong oxidizing property, high catalytic ability, etc. The anatase TiO₂ surface has an important impact on photocatalytic activity, photovoltaic properties, and gas sensor characteristics. Hence, an intensive study of the anatase TiO₂ surface is very significant. The most stable surface of anatase TiO₂ is the (101), followed (001) surfaces^(13,14).

The experimental part of our study aims to evaluate the photodegradation of Alizarine Red Sulfate (ARS) as pollutant substance using TiO₂ anatase as an adsorbance surface and sensitizer with (UV) Light. The effect of initial concentration of (ARS), the PH, and temperature was studied on the rate of photodecomposition of (ARS).

The theoretical part included two calculations, the first was the Equilibrium Geometry optimization at ground state using Density Functional Theory/B3LYP method and Semi-empirical/PM3 method to predict the most stable geometrical formula, the second calculation of the energy at ground state by using Density Functional Theory/EDF2 method to calculate Standard Thermodynamic quantities for pollutant and TiO₂ surface, and to compare between these two studies.

Quantum chemical calculations employing DFT⁽¹⁵⁾ are conducted to investigate the energetics of TiO₂ (101) and (001) anatase surfaces and ARS surface in different mediums (basic, natural, acidic). (101) face represents the most probable anatase surface⁽¹⁶⁾ and (001) surface is considered to be the most stable surface of TiO₂ anatase⁽¹⁷⁾. The anatase TiO₂ cluster contains two Ti atoms and nine O atoms. This cluster approach is a well known and successful approach applied in quantum chemical calculations^(18,19). However, since there may be limitations to the use of small cluster models to represent oxide surfaces where the peripheral oxygen atoms are saturated with hydrogen atoms, the effect of cluster size is also tested by use of semiempirical PM3 calculations. The basis set employed in the DFT calculations was 6-31G** provided in SPARTAN' 08 (Wavefunction Inc.). SPARTAN'08 was used for all the semiempirical and DFT calculations reported, Neutral cluster models of (101) and (001) titanium dioxide anatase surfaces represented by Ti₂O₉H₁₀ but basic cluster models was represented by Ti₂O₉H₉ and acidic cluster models represented by Ti₂O₉H₁₂ for (101) surface whereas (001) surface acidic cluster models was represented by Ti₂O₉H₁₁⁽²⁰⁾ These clusters model geometries are then energetically optimized. A. Cordoba⁽²¹⁾

investigate the possibility of use of Enzymatic treatment by experimental and theoretical study involve molecular modeling through MM2-PM6 by used MOPAC.2009, Vittadini et al⁽²²⁾ indicate that the adsorption of formic acid and sodium formate on the dry and hydrated TiO₂ anatase (101) surfaces a molecular monodentate species hydrogen bonded to a surface two-fold coordinated oxygen is preferred on the dry surface. But on the hydrated surface, HCOOH and HCOONa preferentially form an inner-sphere adsorption complex. Opera et al⁽²³⁾ published that The Ti-O bond lengths for NKX-2398 dye are shorter from (NKX-2311 and C343) dyes, suggesting a strong binding. Selloni et al⁽²⁴⁾ indicated by Molecular Dynamics (MD) calculations adsorption of water be more favorable on the surface TiO₂(101) Anatase These results are agree with the findings of the RedFevn⁽²⁵⁾ But using methods Ab-initio Molecular Orbital Theory Density Functional Theory, Gong et al⁽²⁶⁾ explained TiO₂ (001) anatase surface which contains a lot of Ti5c atoms that give high efficiency on the adsorption molecules such and formic acid, that is agree with M. Vittoria⁽²⁷⁾ which indicate the high activity for TiO₂ (001) anatase surface from Ti5c atoms and large value of (Ti-O-Ti) bond.

2. EXPERIMENTAL DETAILS

Chemicals

All chemicals used during experimental work were of analytical grade and were used as such without purification. Alizarin red S (Riedel de Haën) and HCl (EDH 37%). NaOH (Fluka 99%) were used. Double distilled water was used for the preparation of all types of solution and dilution when required.

Instrumentation

Balance METTLER AE240 and UV/VIS spectrophotometer (PG Instruments Limited U.V- T80) were used in this study.

Standard Solutions

The stock solution of the dye was prepared by dissolving 0.1g/L of dye in water and made up a stock solution in volumetric flask. The concentration of the dye solution was determined spectrophotometrically. Standard solutions of dye were prepared by successive dilution of stock solution.

Experimental Procedure

The photocatalytic degradation of Alizarin red-S for both initial concentration and irradiation sample was determined by UV-Visible Spectrophotometer (PG Instruments Limited U.V- T80). The calibration curve was obtained at λ_{max} =260 nm. The reaction mixture was irradiated with light source UV lamp (PHILIPS-125Watt) at a distance of 15cm from the reaction vessel. Double distilled water was used throughout the experiment. Kinetics of adsorption was determined by analyzing adsorptive uptake of dye from aqueous solution at different time intervals for 15 minutes. In each experiment an accurately weighted amount of TiO₂ was taken in flask in 25 ml dye solution by adjusting pH of the Alizarin red-S dye solution.

3. RESULTS AND DISCUSSION

In order to find the appropriate conditions of adsorbent dose, concentration of dye, pH, and temperature for the adsorption of Alizarin Red S on TiO₂ anatase, various experiments were conducted. The results of these experiments were as followed.

Effect of Adsorbent Dose

The effect of variation in the adsorbent amount on the process adsorption of Alizarin Red S was studied, with different adsorbent amount in the range of (0.1, 0.3, 0.4) g . The results are shown in Fig. 1.

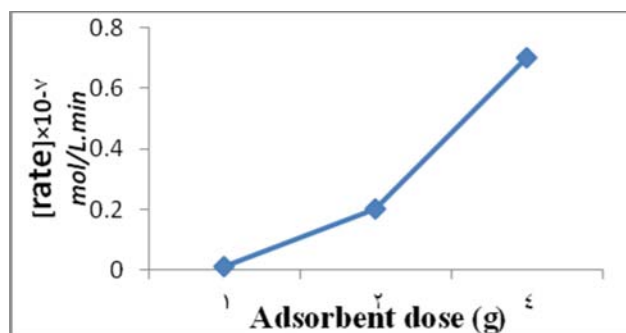


Fig.2 Effect of adsorbent dose on ARS degradation rate. Experimental parameters: Contact time, t (15 min), initial concentration of dye (20 ppm), temperature (30 °C) and solution volume (V) 25 mL .

From Fig.2, an increase in adsorbent dose leads to increase in the percentage of the removed of alizarin. This is caused by an increased in number of available active points with increase in adsorbent dose which mean increase of ($h^+ e^-$) pair number which caused increase in degradation rate⁽¹¹⁾. On the other side when adsorbent dose increase the absorbate photons number well increase that caused increase in degradation rate⁽²⁸⁾.

Effect of Initial Dye Concentration

Initial dye concentration was one of the effective factors on the rate of degradation reaction. The rate of Alizarin Red S degradation was studied as a function initial dye concentration of in the range of 10-30 ppm. The results obtained are present in Fig.2. The rate of (ARS) degradation increases with increasing the initial concentration of the dye TiO₂. It was observed that adsorption yield increased with increase in initial concentration of the dye. Minimum rate was 1×10^{-7} mol/L.min for 10ppm concentration to maximum rate value 1.3×10^{-7} mol/L.min for 30 ppm concentration of dye solution. This may be due to available active sites and increase in the adsorbed molecules which mean molecules involved in the interaction , as an increase in the high initial concentration of the dye.⁽¹⁰⁾

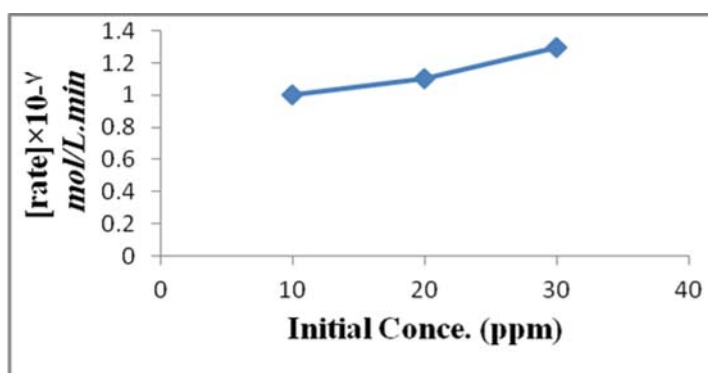


Fig. 3 : Effect of initial concentration of dye on (ARS) degradation rate. Experimental parameters: t(15 min) ,temperature (30 °C) and solution volume (V) 25 mL.

Effect of pH

The pH of the aqueous solution was clearly an important parameter that controls the degradation process. The rate of Alizarin Red S degradation was studied as a function of pH in three mediums. The results obtained were shown in Fig. 3. The minimum degradation rate was 0.3×10^{-7} at pH 5.0 and maximum degradation rate was 0.9×10^{-7} at pH 9 was followed the consequences basic > neutral > acidic .

Because of ARS from pollutants which are hydrolyzed under alkaline conditions may show an increase of reaction rate with increase of pH ⁽²⁹⁾. Finally, the presence of ionic species could affect the degradation process via adsorption of the pollutants, absorption of UV light and reaction with hydroxyl radicals that is leads to increase in degradation rate ^(30,31).

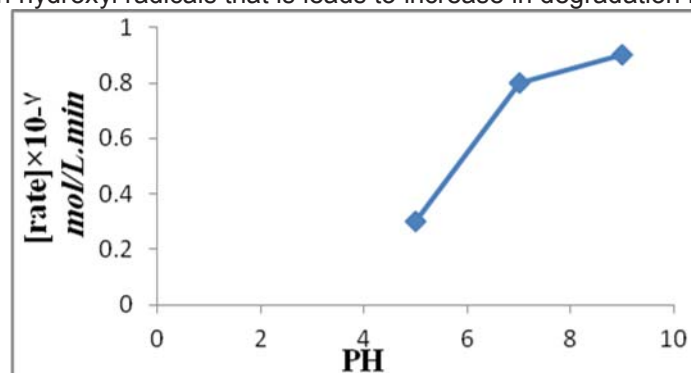


Fig. 4 : Effect of pH on (ARS) degradation rate. Experimental parameters: t(15 min), initial concentration of dye (20ppm), temperature (30 °C) and solution volume (V) 25 mL.

Effect of Temperature

Temperature has an important effect on the rate of adsorption. The rate of Alizarin Red S degradation was studied as a function of temperature in the range of 25-40 °C. The results obtained were present in Fig. 3. It was observed that decrease (ARS) degradation rate with increase in temperature. The minimum degradation rate was at 40 °C and maximum degradation was at 30 °C for 20 ppm initial concentration of dye solution. The decrease in adsorption was obtained at high temperature because molecules move with great speed and less time of interaction was available for dye anions with adsorbent material ^(32,33).

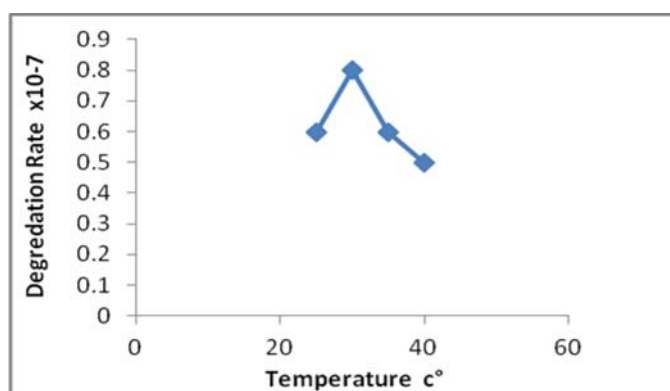


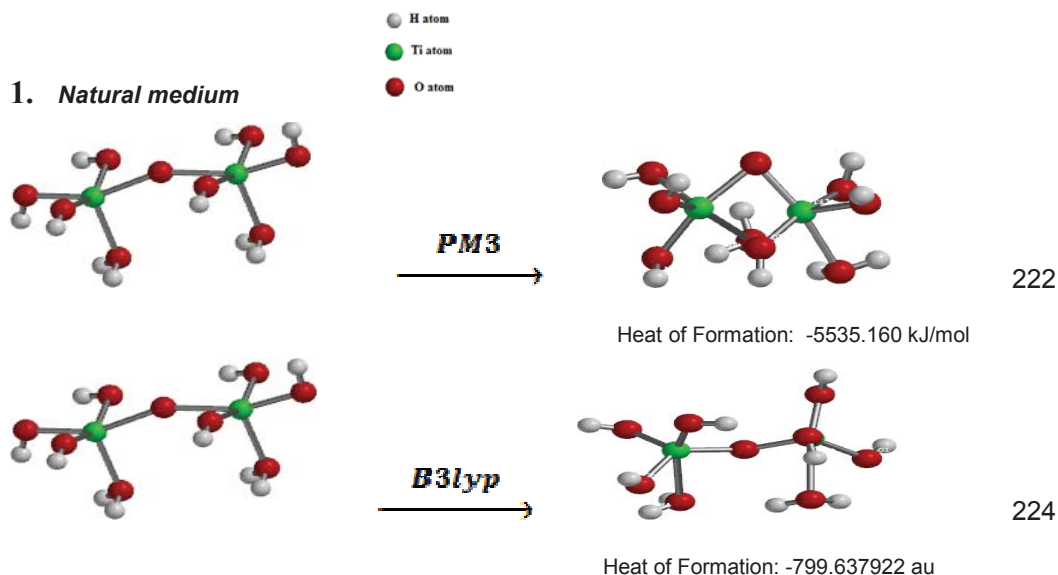
Fig. 5: Effect of temperature on (ARS) degradation rate of Alizarin Red S. Experimental parameters: t(15 min), initial concentration of dye Con. (20ppm), and solution volume (V) 25 mL).

Geometrical Optimization

After selection of the final model for each (ARS) pollutant and TiO_2 surface Anatase (101),(001) the equilibrium geometry was calculated for ground state by (Density Functional Theory / B3LYP) and (PM3 /Semi-Empirical) in the three different mediums (natural, basic, acidic)and calculate the bonds length and angles for the result optimized structure which give us idea of surface stability and reactivity that is resulted from the most stable structure and least reactivity in this mediums. Surface reactivity increase with increase titanium atoms (Ti_{5c}) in the surface and from large value of bond angle (Ti-O-Ti) which from reactivity of O atoms.

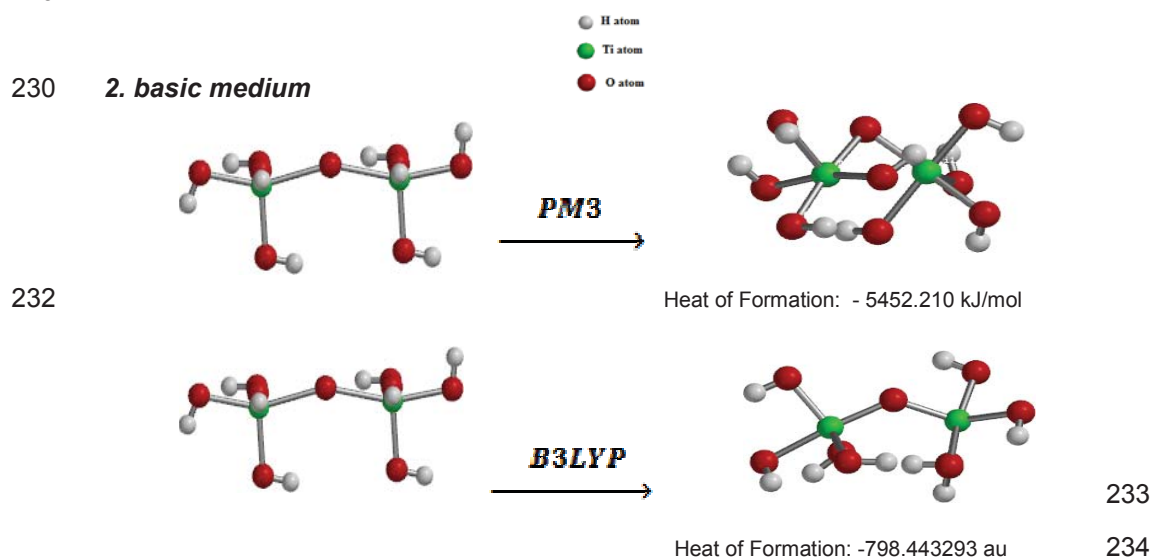
220 TiO_2 anatase (001) Surface Geometrical Optimization

221 1. Natural medium



225 **Figure(1-1):(001)surface Optimized geometrical formula in natural medium .**

230 2. basic medium

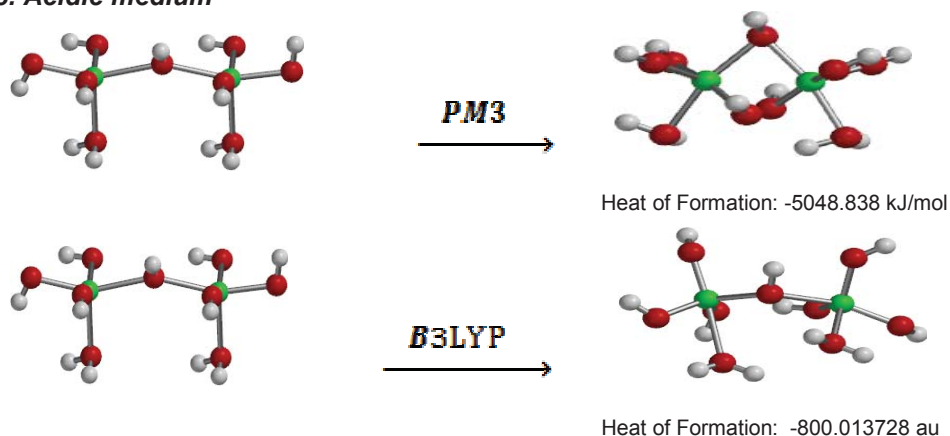


235 **Figure(1-2):(001)surface Optimized geometrical formula in basic medium .**

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3. Acidic medium

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Figure(1-3):(001)surface Optimized geometrical formula in acidic medium .

Table(1-1) bonds length and angle value for (001) surface.

	Medium	Method	Ti _{5C} -O _{2C} (a°)	θ
This Work	Basic	B3LYP	1.831	131.73°
		PM3	1.949	80.70°
This Work	Natural	B3LYP	1.738	129.73°
		PM3	1.950	81.22°
This Work	Acidic	B3LYP	2.138	128.28°
		PM3	2.010	76.78
Ref.34		B3LYP	1.857	-
Ref.36		PBE	-	146°

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θ : value of (Ti-O-Ti) bond

Ti_{5C} : fivefold Titanium atom

O_{2C} : twofold oxygen atom

249 From table (1-1) value of (θ) angle showed that the (001) surface has most value from
250 others that is mean highest reactivity in basic medium and showed that the method
251 DFT/B3LYP give the best results and confirmed with The practical methods that the basic
252 medium give the highest rate of decomposition of (ARS) followed the consequences:

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basic >neutral >acidic

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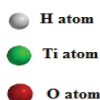
whereas (Ti_{5C} -O_{2C}) bond lengths in acidic medium has most value from others that is mean
non favorite medium for TiO₂ anatase (001) surface.

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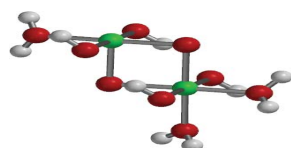
TiO₂ anatase (101) Surface Geometrical Optimization

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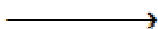
1. natural surface



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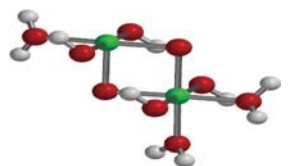
PM3



Heat of Formation: -5441.139 kJ/mol

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B3lyp



Heat of Formation: -799.646018 au

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figure(1-4):(101)surface Optimized geometrical formula in natural medium .

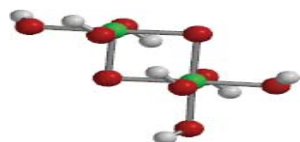
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2. basic medium

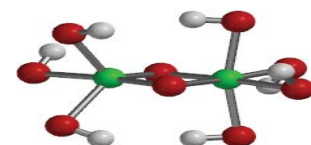
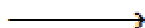
● H atom

● Ti atom

● O atom



PM3



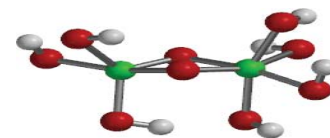
Heat of Formation: - 4798.750 kJ/mol

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B3LYP



Heat of Formation: -797.608001 au

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figure(1-5):(101)surface Optimized geometrical formula in basic medium .

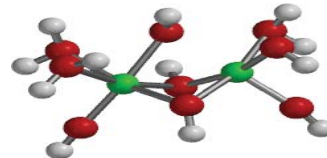
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3. acidic medium



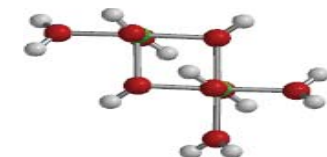
PM3



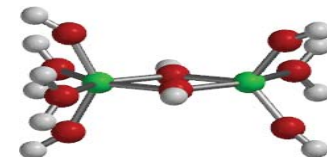
Heat of Formation: -4066.368 kJ/mol

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B3LYP



Heat of Formation: -800.256054 au

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figure(1-6):(101)surface Optimized geometrical formula in acidic medium.

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Table(1-2) bonds length and angle value for (101) surface

	Medium	Method	Ti _{5C} -O _{2C}	Ti _{6C} -O _{2C}	θ
This Work	Basic	B3LYP	1.953	1.831	99.13
		PM3	1.896	1.981	83.09
This Work	Natural	B3LYP	1.873	1.823	96.71
		PM3	1.917	1.938	80.36
This Work	Acidic	B3LYP	2.004	1.979	109.38°
		PM3	2.014	2.027	75.70°
Ref.34		B3LYP	1.966	1.937	-
Ref.35		B3LYP	1.830	1.850	-
Ref.37		GGA	1.836	1.849	-
Ref.36		PBE	1.942	-	102

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θ: value of (Ti-O-Ti) bond

Ti_{5C}: fivefold Titanium atom

Ti_{6C}: sixfold Titanium atom

O_{2C}: twofold oxygen atom

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Each methods PM3 and B3LYP showed that the θ value for TiO₂ anatase (101) surface in the basic medium is more than that's in natural medium which refers to high reactivity for (101) surface in basic medium but in acidic medium values varied where(B3LYP) method give large value of θ which refers to high reactivity from the basic and natural mediums that doesn't agree with the experimental results while PM3 method showed small value of θ which refers to low reactivity less than in basic and natural mediums which agree with the experimental results.

Comparing between TiO₂ anatase (001) and (101) surfaces by bonds length and angle value, we find that the (001) surface was more reactive than (101) surface. Our results agree with the previous findings reported in different studies^(38,39,26).

Table(1-3): heat of formation for (001),(101) surface in different media .

	Method	Basic	Neutral	Acidic	Energy Unit
001	PM3	-5452.21	-5535.16	-5089.44	kJ/mol
	B3LYP	-798.443	-799.639	-800.013	au
101	PM3	4798.750-	-5441.139	-4066.36	kJ/mol
	B3LYP	-797.6	-799.646	-800.256	au

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PM3 methods in table (1-3) showed that both TiO₂ anatase(001) and (101) surfaces has less heat of formation in neutral medium which means each of them had highest stability and lowest reactivity in neutral medium, whereas in acidic medium both surfaces have high heat of formation which means lowest stability and highest reactivity in acidic medium.

B3LYP methods showed that both TiO₂ anatase (001) and (101) surfaces have high heat of formation in basic medium which means each of them had lowest stability and highest reactivity in basic medium, whereas in acidic medium both surface have lowest heat of formation which means highest stability and lowest reactivity ,that's agree with the experimental results which indicate ARS degradation rate is the most in basic medium and this was very clear in 101 surface where the highest heat of formation was observed.

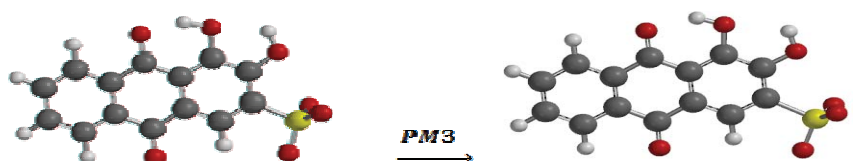
On the other hand comparing between calculated heat of formation by B3LYP method for both TiO₂ surfaces showed that 101 surface has highest heat of formation which means low stability and high reactivity than 001 surface. This result agree with previous studies^(24,25)

Geometrical Optimization) Alizarin Red Sulfate (ARS)

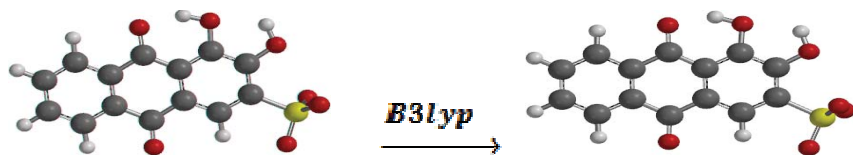
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1. Natural medium





Heat of Formation: -1081.805 kJ/mol

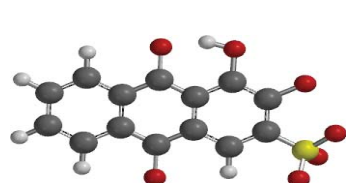


Heat of Formation: -1462.497676 au

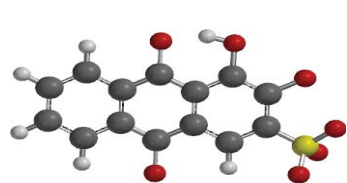
figure(1-7):ARS pollutant Optimized structure in natural medium by(PM3)(B3LYP) .



2. Basic medium



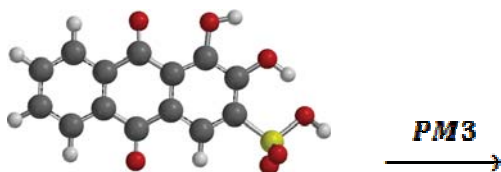
Heat of Formation: -996.948 kJ/mol



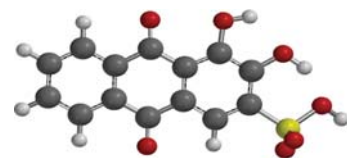
Heat of Formation: -1461.857569 au

figure(1-8):ARS pollutant Optimized structure in basic medium by(PM3)(B3LYP).

3. Acidic medium



Heat of Formation: -902.387 kJ/mol



Heat of Formation: -1462.995253 au

figure(1-9):ARS pollutant Optimized structure in acidic medium by(PM3)(B3LYP).

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335 **Table(1-4) : heat of formation for ARS pollutant in different medium .**

	<i>Method</i>	<i>Basic</i>	<i>Natural</i>	<i>Acidic</i>	<i>Energy Unit</i>
ARS	<i>PM3</i>	996.948-	-1081.805	-902.387	<i>kJ/mol</i>
	<i>B3LYP</i>	-1461.857	-1462.497	-1462.995	<i>au</i>

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337 Heat of formation values in table (1-4) was calculated by B3LYP showed that's ARS
 338 pollutant has highest heat of formation in basic medium fallowed the consequences basic
 339 > neutral > acidic whereas PM3 method showed that's highest heat of formation in acidic
 340 medium fallowed the consequences acidic > basic > neutral that's indicate B3LYP
 341 methods give good results which is agree with the experimental results .

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344 **Theoretical Calculations for Surface and Pollutant**

345 The thermodynamic calculation was performed depending on the energy calculations at
 346 ground state by using Density Functional Theory/EDF2 method to calculate Standard
 347 Thermodynamic quantities for pollutant and TiO₂ surface.

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349 **Tabel(1-5):Standard Thermodynamic quantities for (001),(101) surface in different medium at**
 350 **298.15 K and 1.00 atm**

	<i>Medium</i>	<i>ZPE</i> <i>kJ/mol</i>	<i>Enthalpy</i> <i>kJ/mol</i>	<i>Entropy</i> <i>J/mol.K</i>	<i>Cv</i> <i>J/mol.K</i>	<i>Gibbs free</i> <i>energy G (kJ/mol)</i>
001	Basic	254.6410	344.2600	528.5494	254.6410	186.916
	Natural	361.9670	409.9561	537.8440	266.3754	249.924
	Acidic	388.7331	434.5602	522.5073	257.6182	279.004
101	Basic	268.0738	313.1480	517.0250	250.3672	159.082
	Natural	361.6021	414.0866	574.2387	284.7062	243.028
	Acidic	412.2940	470.5403	606.5702	317.5846	289.952

351 *C_v* : heat capacity

352 *ZPE*: Zero point energy

353 **Tabel(1-5): Standard Thermodynamic quantities for ARS pollutant in different medium at 298.15**
 354 **K and 1.00 atm**

	<i>system</i>	<i>ZPE</i> <i>kJ/mol</i>	<i>Enthalpy</i> <i>kJ/mol</i>	<i>Entropy</i> <i>J/mol.K</i>	<i>Cv</i> <i>J/mol.K</i>	<i>Gibbs free</i> <i>energy G kJ/mol</i>
ARS/EDF2	<i>Basic</i>	467.4259	511.7977	527.8044	268.1554	354.4328
ARS/EDF2	<i>Natural</i>	502.9677	549.7038	546.4649	279.8043	386.7753
ARS/EDF2	<i>Acidic</i>	532.9485	577.9819	532.0865	274.4549	419.3403

355

356 The theoretical thermodynamic values for both (101),(001) surfaces and ARS pollutant in
 357 different mediums in tables(1-4),(1-5) the increase of this values Means an increase of
 358 stability and decrease of reactivity fallowed the consequences :

359 Increase reactivity ← basic > neutral > acidic → Increase stability

360 This results agreed with the experimental results that showed the highest reactivity and
 361 lowest stability for each pollutant (ARS) and TiO₂ surface in the basic medium .

362

363 **4. CONCLUSION**

364

365 The results from theoretical study agreed with the results obtained in the experimental
 366 study. Results from the experimental study revealed that the photodegradation reaction

from pecideu first order and exothermic reaction and optimum parameters for removal of this dye from aqueous media were (0.16 kg/l, 20ppm, temp.30c, pH9). Theoretical study DFT/B3LYP method showed better results than DFT/PM3 to predict the most stable geometrical formula, and TiO₂ anatase (001) surface was more active than TiO₂ anatase (101) surface.

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COMPETING INTERESTS

Authors have declared that no competing interests exist.

AUTHORS' CONTRIBUTIONS

Author THT performed the experiments, statistical analysis and wrote the first draft of the manuscript. Author RRA designed the study, supervise the experiments, and wrote the

480 protocol. Author MAA supervise the theoretical chemistry part and managed the analyses of
481 the study.

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485 **DEFINITIONS, ACRONYMS, ABBREVIATIONS**

486 Here is the Definitions section. This is an optional section.

487 **Term:** Definition for the term

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