



SDI Review Form 1.6

Journal Name:	American Chemical Science Journal
Manuscript Number:	2014_ACSj_11703
Title of the Manuscript:	The Electron Transfer Reaction of 3,7-Bis(dimethylamino)phenazothionium Chloride with Benzenethiol in Aqueous Nitric Acid: A Mechanistic Approach
Type of the Article	Original Research Article

General guideline for Peer Review process:

This journal's peer review policy states that **NO** manuscript should be rejected only on the basis of '**lack of Novelty**', provided the manuscript is scientifically robust and technically sound.

To know the complete guideline for Peer Review process, reviewers are requested to visit this link:

(<http://www.sciencedomain.org/page.php?id=sdi-general-editorial-policy#Peer-Review-Guideline>)



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PART 1: Review Comments

	Reviewer's comment	Author's comment (if agreed with reviewer, correct the manuscript and highlight that part in the manuscript. It is mandatory that authors should write his/her feedback here)
Compulsory REVISION comments	1. In the supposed mechanism of the process, presented in the paper, there is no electron transfer step. It is quite possible, however, that just this step will be slowed-down one, i.e. determine the kinetics of the process. By the way, the authors of the paper point to electron transfer in the paper and table titles. 2. It is not clear from text how the equilibrium concentration of hydrogen ions $[H^+]$ was determined. Was it measured with an instrument or simply as an added amount of acid? Therefore it is difficult to understand from what values the ionic strength was calculated. For example, Fig 2 or 3: is the concentration of hydrogen ions measured, calculated or added? 3. How was the equilibrium concentration $[HCOO^-]$ calculated, and what dissociation constant was used? 4. In the presented mechanism of the process, equilibrium constants were written, but the state of ions depending on pH has not been checked.	Some of the comments here have been taken care of in our earlier revised article. The concentration of the hydrogen ion was calculated so also was the ionic strength of the reaction medium



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<u>Minor</u> REVISION comments	1. Why was the temperature 29°C and not the standard temperature 25°C chosen? 2. The tables and figures contain the same results, they can be partially reduced.	This is the laboratory temperature at the time of this investigation.
<u>Optional/General</u> comments	It would be simpler to write: the measurements were made in 1 M NaNO_3 .	This is not necessary as mol dm^{-3} is more preferable than M in modern Chemistry