



SDI FINAL EVALUATION FORM 1.1

PART 1:

Journal Name:	American Chemical Science Journal
Manuscript Number:	2014_ACSj_14425
Title of the Manuscript:	Adsorption Isotherm Studies of Cd(II) and Pb(II) ions from Aqueous Solutions by Bamboo-Based Activated Charcoal and Bamboo Dust.
Type of the Article	Original Research Article

PART 2:

FINAL EVALUATOR'S comments on revised paper (if any)	Authors' response to final evaluator's comments
- Why have you selected this mass (1.0 g) and pH (6) for doing your experiments? Explain completely - Remove unnecessary metals from Table 4 (such as Ni(II), (Zn(II) and Cu(II)) - Check x and y axis of figure 10, plot of q_e vs $\ln C_e$	In the previous research work carried out by Ademiluyi and Ujile (2013) on Kinetics of batch adsorption of iron (II) ions from aqueous solution using activated carbon from Nigerian Bamboo, it was carried out using an adsorbent mass (dose) of 1 g/L and a reference to that had been stated in line 111-113 of the manuscript and included in the reference section of the present work. Also a preliminary results on influence of pH on the removal of the two metal ions Pb(II) and Cd(II) revealed that the optimum pH for adsorption of the two metal ions by both adsorbents (BBAC and BD) occurred at a pH 6.0 and this has been included in lines 168-174 of this work. The unnecessary metals from Table 4 have been removed. The proper titles for x and y axis of Fig. 10 (now Fig.12) had been given.